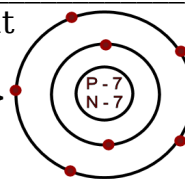


Name _____ Period _____

Bohr and Lewis Models, and Molecular Weight

1. Create Bohr models of elements 1-6. Element 7 shown ==>

12 points



2. Create Lewis structures for elements 1-20. (20 points)

<http://www.green-planet-solar-energy.com/electron-dot-structure.html>

H He Li Be B

C N O F Ne

Na Mg Al Si P

S Cl Ar K Ca

32 points

3. What is the rule on capitalization for elements? (1 point)

4. Find the mass of the following in amu (atomic mass units): (round to tenths)

_____ Na	_____ Cl	_____ NaCl
_____ H	_____ O	_____ H ₂ O
_____ C	_____ F	_____ CH ₄ O
_____ HCl	_____ HF	_____ NaOH

(12 points)

5. Draw the electron configurations of Na and Cl and demonstrate the bond.

(1 point)

6. Write the formula (other than NaCl) of a single bonded ionic molecule.

7. Write the formula for a double bonded ionic molecule.

(1 point)

(1 point)

8. Define: What is a covalent bond? (1 point)

9. Give two examples of covalent bonds. (2 points)

10. Water is a polar covalent bond. What is a polar covalent bond? (1 point)

11. Draw a water molecule using a Lewis dot structure

12. Draw a Bohr model of water

(1 point)

(1 point)